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Solvent control of charge transfer excited state relaxation pathways in $[\text{Fe}(\text{2,2'}\text{-bipyridine})(\text{CN})_4]^{2-}$

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The excited state dynamics of solvated $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$, where bpy = 2,2'-bipyridine, show significant sensitivity to the solvent Lewis acidity. Using a combination of optical absorption and X-ray emission transient spectroscopies, we have previously shown that the metal to ligand charge transfer (MLCT) excited state of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ has a 19 picosecond lifetime and no discernable contribution from metal centered (MC) states in weak Lewis acid solvents, such as dimethyl sulfoxide and acetonitrile.^{1,2} In the present work, we use the same combination of spectroscopic techniques to measure the MLCT excited state relaxation dynamics of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in water, a strong Lewis acid solvent. The charge-transfer excited state is now found to decay in less than 100 femtoseconds, forming a quasi-stable metal centered excited state with a 13 picosecond lifetime. We find that this MC excited state has triplet (³MC) character, unlike other reported six-coordinate Fe(II)-centered coordination compounds, which form MC quintet (⁵MC) states. The solvent dependent changes in excited state non-radiative relaxation for $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ allows us to infer the influence of the solvent on the electronic structure of the complex. Furthermore, the robust characterization of the dynamics and optical spectral signatures of the isolated ³MC intermediate provides a strong foundation for identifying ³MC intermediates in the electronic excited state relaxation mechanisms of similar Fe-centered systems being developed for solar applications.

Introduction

Inorganic complexes are attractive candidates for solar energy applications due to their tunable electronic properties, which

can be varied synthetically by changing the metal atom and ligand composition or structure. Molecular photosensitizers based on 4d and 5d transition metals, such as Ru and Ir, have been successfully implemented in photovoltaic applications

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due to their intense absorption, long-lived charge transfer excited states, and their ability to undergo reversible redox processes.³ However, their cost and scarcity render them impractical for large-scale development.

Transferring the functionality of 4d and 5d transition metal complexes to the more abundant 3d transition metal centered systems motivates wide-ranging investigations targeting solar energy applications.⁴ Solar energy applications benefit from long-lived metal-to-ligand charge-transfer (MLCT) excited states, but the majority of optically generated MLCT excited states in 3d transition metal complexes relax on the sub-ps timescale. A central challenge to extending the charge transfer excited state lifetimes in light harvesting complexes based on 3d transition metals is the weaker ligand field splitting of 3d complexes, relative to their 4d or 5d counterparts. The smaller ligand field splitting leads to ligand field excited states with lower energies than the optically bright charge transfer excited states that very efficiently quench the MLCT excited states. This deactivation reduces the MLCT lifetime of 3d transition metal complexes by orders of magnitude compared to their 4d and 5d analogs. For instance, ruthenium(II)-centered polypyridyl compounds are characterized by MLCT states with hundreds of nanoseconds to microseconds lifetimes,^{5–9} the MLCT state of similar complexes with an isoelectronic iron(II) center undergoes spin crossover to a metal centered quintet (⁵MC) state within hundreds of femtoseconds.^{10–24}

Significant progress in the design of iron-centered molecular systems with extended MLCT lifetimes targeted at solar energy applications has been made recently and demonstrates the potential of suppressing internal conversion and intersystem crossing.^{2,25–30} The synthetic strategy pursued with the most success for octahedral iron(II) complexes has been ligand selection for increased ligand-to-metal σ -donation, beyond that of typical polypyridyl ligands. Increased σ -donation leads to larger ligand field splitting and destabilization of the MC excited states relative to MLCT excited states. A range of strong σ donating N-heterocyclic carbene ligands have provided Fe-complexes with MLCT lifetimes ranging from ~ 300 fs to ~ 30 ps,^{2,25,27,30,31} long enough to allow transfer of the ligand-localized electron for photovoltaic applications.^{27,28}

The metal centered triplet (³MC) excited state has recently been suggested to play a critical role in mediating the MLCT decay pathway in the cases of both Fe(II)-centered spin crossover complexes (short-lived, ~ 100 fs, MLCT states)^{21,22} and MLCT complexes (long-lived, ~ 20 ps, MLCT states),^{2,26,29,30} and has additionally been observed as a precursor for ligand dissociation following excitation of $[\text{Fe}(\text{CN})_6]$.^{4–32} Although still under debate,¹⁸ much experimental and computational evidence supports the presence of a very short lived ³MC transient state in the spin state transition dynamics to the relatively long lived ⁵MC state.^{20–22,33,34} In the cases where MLCT lifetimes are extended to the picosecond time scale, they decay without well characterized intermediates, thus preventing experimental identification of decay-mediating states. Quantum chemical and dynamics calculations for several such complexes do suggest that the MLCT lifetime is strongly influenced by the

relative positions of the potential curves of the MLCT and ³MC excited states.^{2,26,29,35–38} Although the important role of the ³MC state in facilitating the excited state decay of hexacoordinated Fe(II) complexes has been suggested, a metastable ³MC state has yet to be cleanly isolated for these complexes. Therefore, it is difficult to assess the capacity of experimental methods to identify the role of ³MC intermediates in the excited state dynamics of Fe-based systems.

A strong solvatochromic effect has been well-established for Fe- and Ru-centered cyano-polypyridyl complexes,^{39–44} and has been ascribed to the interaction of high Lewis acidity solvents (such as water) with the N lone pair of the CN^- ligand. Here, we exploit the solvent-dependent electronic structure of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$, illustrated in Fig. 1, to vary the relative energies of the MLCT and MC states with the intent of generating a metastable ³MC. The choice of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ was motivated by previous transient optical measurements that demonstrate the loss of characteristic MLCT features from the excited state spectrum upon changing the solvent Lewis acidity.⁴⁵ This study clearly demonstrated the solvent dependence of the non-radiative relaxation in this complex, but did not have the time resolution to determine relaxation rates nor the ability to determine the MC excited state intermediate formed in water.⁴⁵ The use of such a chemically amenable system allows us to develop our current understanding of how metal-ligand bonding controls the internal conversion and intersystem crossing pathways in 3d transition metal complexes. In this regard, the solvatochromic $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ mixed ligand complex provides an excellent model for addressing these fundamental questions in physical chemistry, as the degree of metal-ligand σ - and π -bonding interactions can be varied with the solvent. Steady-state measurements have shown this interaction to shift electron density on the cyanide ligand and increase the metal-to-ligand π

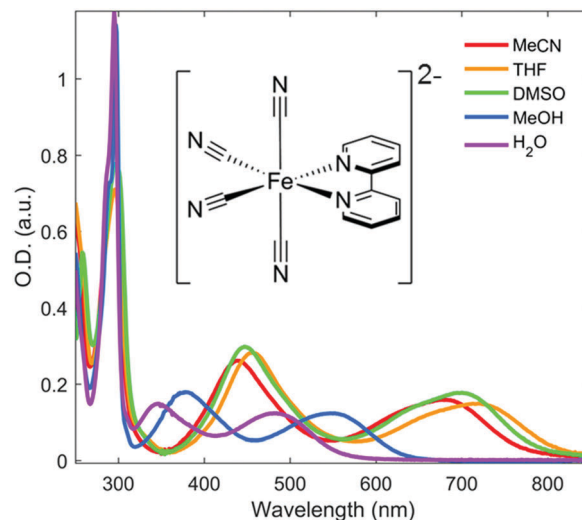


Fig. 1 Solvatochromism of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$. The inset shows the $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ system, and the rest of the panel shows its optical absorption spectrum in a series of solvents with varying Lewis acidity; acetonitrile (MeCN) tetrahydrofuran (THF), dimethyl sulfoxide (DMSO), methanol (MeOH), and water (H_2O).

back-donation for the water-solvated complex.^{39,40,46} The highest occupied molecular orbitals (HOMO) of mixed Fe d (t_{2g}) and cyano π^* character are therefore stabilized relative to the lowest unoccupied molecular orbitals (LUMO) of bpy-ligand π^* character, resulting in a net destabilization of the respective MLCT excited state.⁴⁷ Experimental investigation of the solvent influence on the MC levels is significantly more scarce, but one study finds the trend of solvent-dependent excited state lifetimes for Ru cyanopolypyridyl complexes can be qualitatively reproduced under the assumption that the MC energy levels are unchanged relative to the ground state.⁴⁴

To characterize the excited state dynamics of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in water we have combined UV-visible pump-probe spectroscopy measuring the transient optical absorption (TA) with transient K-edge X-ray emission spectroscopy (XES), a combination which have delivered key insights in the dynamics of 3d transition metal centered systems before.^{1,21,48–52} The signal recorded by TA arises from dipole-allowed transitions in the optical regime, rendering it sensitive to charge transfer excited states involving ligand-localized excited electrons. XES monitors the fluorescence emitted during the 2p-to-1s or 3p-to-1s decay following 1s core-hole ionization of the Fe metal center. The fluorescence intensity and spectral features are sensitive to the total spin moment on the iron center, and therefore, this method can distinguish MC excited states differing in Fe spin moments, such as ^3MC and ^5MC states. The combination of the two techniques enables full characterization of the excited state relaxation pathway of 3d transition metal complexes. The present study has unambiguously identified a metastable ^3MC excited state with the transient XES spectrum. This identification, in combination with the TA measurements enables the assignment of transient absorption features in the UV-visible spectrum to the ^3MC excited state that should prove valuable to the interpretation of electronic excited state relaxation dynamics in other Fe(II) systems.

We report the excited state dynamics of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ solvated in water and identify a temporally isolated ^3MC intermediate for the first time. In contrast, we have previously demonstrated a long lived MLCT excited state (19 ps) for $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in weak Lewis acidic solvents (dimethyl sulfoxide (DMSO) and acetonitrile (MeCN)), with no discernable contribution from metal centered states.¹ In the present work, we shift the bpy-localized MLCT state towards higher energy by using a high Lewis acidity solvent results in a faster MLCT-to-MC transition and slower MC-to-GS transition. The identification of a metastable ^3MC state provides further support for its role in deactivating MLCT excited states in Fe(II) complexes, and expands on the previously reported studies on Ru-centered systems which were unable to distinguish the influence of the solvent on the MC state energy levels.⁴⁴

Experimental methods

X-ray emission

The time-resolved XES measurements were conducted at the X-ray Pump-Probe (XPP)⁵³ end station at the Linac Coherent

Light Source (LCLS). An aqueous solution of 55 mM $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ was pumped through a 50 μm diameter nozzle producing a cylindrical liquid jet. The sample was excited with 400 nm optical laser pulses of 45 fs duration, 120 μm focus diameter (FWHM), and 12.5 μJ per pulse, delivering a pulse fluence of $3 \times 10^{12} \text{ W cm}^{-2}$. The sample was probed by 8.5 keV X-ray laser pulses of ~ 30 fs duration. The Fe 3p-1s ($\text{K}\beta$) fluorescence XES signal was detected on a 140k Cornell-SLAC Pixel Array Detector (CSPAD) area detector⁵⁴ located above the liquid jet using four dispersive Ge(620) crystal analyzers with a central Bragg angle of 79.1 degrees.⁵⁵ The Fe 2p-1s ($\text{K}\alpha$) fluorescence XES signal was detected on a second 140k CSPAD detector placed behind the sample, using a spherically bent Ge(440) crystal analyzer, set to a Bragg angle corresponding to the maximum signal of the iron $\text{K}\alpha$ fluorescence of 6404 eV (75.4 degrees).

The full 2D images of the XES detectors were read out for each pump-probe event, normalized, and corrected as described in the ESI.[†] Difference images were constructed by subtracting a reference signal recorded from optical laser-off probe events for every seventh X-ray laser pulse event throughout the data collection. The difference images of each pump-probe event were then sorted into 250 individual time bins (400 shots per bin) according to their individually recorded time delay (see the ESI,[†] for a description of the timing tool diagnostic). The $\text{K}\alpha$ difference images for each time bin were averaged and integrated, resulting in the kinetic trace presented in Fig. 5, while the $\text{K}\beta$ difference images of each time bin were averaged and integrated along the nondispersive detector axis, resulting in the difference spectra presented in Fig. 6A.

The static $\text{K}\alpha$ XES spectra of reference compounds were measured at the Stanford Synchrotron Radiation Lightsource (SSRL) beamline 6-2. Reference compounds were obtained from Sigma-Aldrich and used as-is. Samples were measured as powders (covered by Kapton tape) in a cryostat at 10 K. Samples were excited with monochromatic incident X-rays at 7300 eV (double-crystal Si(311) monochromator, 0.2 eV FWHM bandwidth) and Fe $\text{K}\alpha$ fluorescence was detected with a Rowland geometry spectrometer utilizing (440) Bragg reflection from bent Ge crystals with a 1 m radius of curvature (energy resolution 0.6 eV FWHM).⁵⁶ X-rays were detected with a Si drift detector and a correction for detector nonlinearity was applied. The monochromator energy was calibrated with an Fe foil and the spectrometer calibration and transmission correction was done using elastic scattering. Each sample was checked for X-ray induced damage by measuring the Fe X-ray absorption near-edge structure (XANES) spectra at various incident fluxes using filters installed in the beamline. Each sample was measured multiple times at different positions to reduce exposure.

Optical transient absorption

Femtosecond time-resolved UV-visible transient absorption (TA) measurements were conducted on $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ samples in H_2O (7 mM) and DMSO (2 mM), prepared directly before the experiments. Liquid sheet jets with 100 μm and 200 μm thicknesses were used to deliver the sample into the pump-probe overlap region, respectively. Continuous flow of the jet

ensured that any accumulation of photo-damaged sample was avoided.

Experiments were carried out using an amplified Ti:sapphire laser system (Coherent Mantis with Coherent Legend Elite Duo) with a 5 kHz repetition rate, 800 nm central wavelength, 2 mJ pulse energy and 40 fs FWHM pulse duration. A portion of the laser pumped an optical parametric amplifier (Spectra-Physics OPA-800C) to generate near IR signal and idler *via* difference frequency mixing. The output of the OPA was used to generate 500 nm pump pulses *via* sum-frequency generation of the signal with 800 nm light, and to generate 700 nm pump pulses *via* frequency doubling of the signal in a BBO crystal. The pump pulse was directed to the sample through a delay stage, a 2.5 kHz chopper, and a lens, resulting in pump pulses with 100 μm focus diameter (FWHM), 50 fs duration, and 1.6 μJ pulse energy providing a pump fluence of $4 \times 10^{11} \text{ W cm}^{-2}$. The pump was overlapped with a white light probe pulse (*via* supercontinuum generation in 4 mm of CaF_2) at the sample position. The probe was transmitted through the sample and imaged on a spectrometer (Horiba Jobin Yvon iHR320, grating 150 grooves per mm). The probe spectrum was recorded at 5 kHz with a NMOS linear image sensor (Hamamatsu, S8380-512Q) simultaneously over the whole 300–700 nm spectral range. The differential absorbance (ΔA) was calculated as $\Delta A = \log(I_{\text{off}}/I_{\text{on}})$, where I_{on} and I_{off} are the pumped and unpumped intensity, respectively. Overall time-resolution of the experiment was approximately 100 fs.

Results and discussion

Steady-state and transient optical absorption spectroscopy

The steady-state optical absorption spectrum of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in a series of solvents is presented in Fig. 1. In weak Lewis acid solvents (acetonitrile (MeCN, red), tetrahydrofuran (THF, orange), and dimethyl sulfoxide (DMSO, green)) the visible spectrum is

characterized by two absorption features at roughly 450 nm and 700 nm assigned to MLCT excitations from orbitals of mixed Fe d and cyano π and π^* character to π^* orbitals of the bipyridine ligand.⁵⁷ For high Lewis acidity solvents, such as the hydrogen-bonding methanol (MeOH, blue) and H_2O (purple), these absorption features blue-shift. This solvatochromic effect has been extensively mapped for cyano-polypyridyl Fe and Ru complexes and shown to be linear, both with respect to the solvent acceptor number (a measure of the Lewis acidity) and the number of cyano ligands present.^{39,40,42,57} This makes $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in solution an archetypical system to characterize the influence of the solvent Lewis acidity on the properties of these complexes.

The blue-shift of the MLCT absorption band clearly shows that increasing solvent Lewis acidity strongly destabilizes the MLCT state with respect to the ground state. To investigate how the solvent influences the energy levels of the MC states, and to determine how this modifies the MLCT deactivation pathways, we turn to the solvent dependence of the excited state relaxation dynamics.

Fig. 2 shows the solvent-dependent differential absorption spectra of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ obtained after excitation of the lowest energy MLCT transition (700 nm in DMSO, 500 nm in water). In DMSO (Fig. 2A), the differential spectra at all time delays are characterized by the ground state bleach (GSB), overlaid with a broad excited state absorption (ESA) feature everywhere below 600 nm, peaking at 370 nm and 525 nm. In bipyridine-containing complexes, the strong 370 nm ESA feature is typically associated with an excited electronic state having an electron localized on the pyridyl ligand.^{1,2,23,58,59} In previous work, we have complemented such TA measurements by K β XES, determining that the excited state dynamics are well-described by a 19 ps single-exponential decay of a bipyridine-localized MLCT state.¹ In water (Fig. 2B), the difference spectra at very early time delays (~ 100 fs) can also be described by the GSB overlaid by a similar ESA feature everywhere below

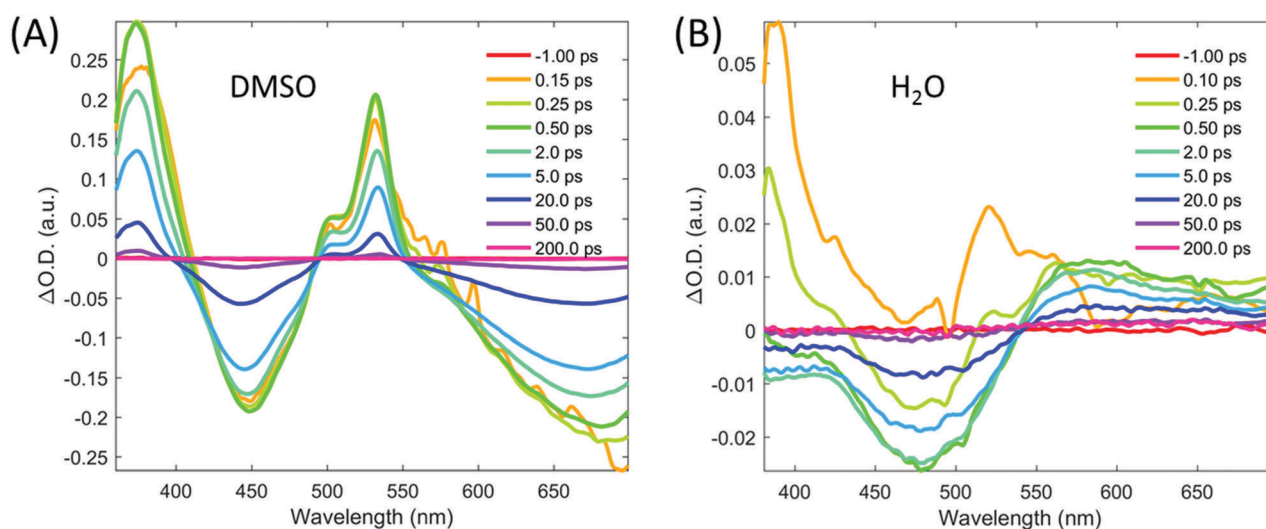


Fig. 2 Transient absorption spectra of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ at selected time delays in (A) DMSO with 700 nm excitation and (B) water with 500 nm excitation.

600 nm, peaking at 370 nm and 525 nm. However, within half a picosecond, the initial ESA features decay completely. Differential spectra at time delays longer than 0.5 ps are characterized by the GSB and a broad ESA feature above 600 nm. These features decay together over the following 50 ps as the ground state spectrum recovers. For $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in water, the ultrafast appearance and decay of the 370 nm ESA feature without the decay of the ground state bleach indicates that the excitation of the MLCT band initially populates a very short-lived bipyridine-localized MLCT state, which decays to a secondary electronic excited state before returning to the ground state on the tens of picosecond time scale. As described in the ESI,[†] high excitation fluencies lead to a longer lived absorption feature centered at 680 nm characteristic of the solvated electron⁶⁰ arising from multiphoton ionization.

From the optical data presented in Fig. 2B, the excited state cascade in aqueous solution can be formulated as an instantaneous generation of an MLCT state and a small fraction (<0.01) of photoionization products. The MLCT state undergoes ultrafast conversion to a second excited state which decays on the picosecond timescale, while the photoionization product remains. When this three-state model is implemented in a global analysis framework and applied to the data, the species associated spectra (SAS) shown in Fig. 3 are identified. SAS1, which we associate with the MLCT state, decays with a 0.17 ± 0.03 ps lifetime. SAS2, which we associate with the second excited electronic state, decays with a lifetime of 12.7 ± 0.4 ps. SAS3, which we associate with the solvated electrons remains within the 1 ns time window of these measurements.

To facilitate the direct comparison of the excited state dynamics measured with TA at low fluence to those measured by XES at higher fluence, we have applied the same analysis to additional TA measurements recorded at higher fluence, which exhibit a significant contribution from solvated electrons. The transient spectra and resulting SAS are shown in the ESI.[†]

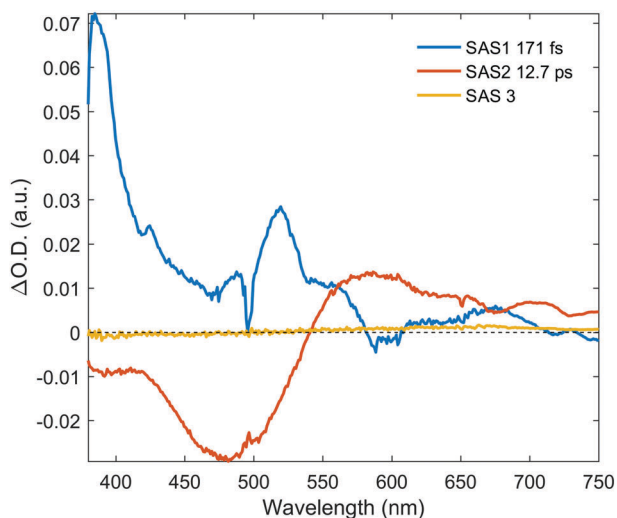


Fig. 3 Species associated spectra (SAS) from a global analysis of the transient absorption data for $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ dissolved in water and shown in Fig. 2B.

The spectral features of the photoexcited states (SAS1 and SAS2) are virtually identical between the high and the low fluence measurements, and the variation in the extracted relaxation dynamics lies within the uncertainties of the measurements. This demonstrates that the decay pathway of the single-photon MLCT non-radiative relaxation is independent of the excitation conditions over a large range of pump fluence.

Assigning the nature of the second electronic excited state associated with SAS2 in the global analysis is difficult based on optical TA data alone. The broad and relatively weak >600 nm ESA is the only significant spectral feature outside the ground state bleach. This ESA lacks significant or defining features that might otherwise allow assignment of the associated species. The short lifetime and complete ground state recovery rules out ionization and degradation products as candidates for SAS2. The lack of low-lying unoccupied ligand-centered electronic states (other than the bipyridine π^* states) suggests that SAS2 describes a secondary excited state with MC character. To unambiguously determine the nature of this intermediate, we turn to the transient $K\alpha$ and $K\beta$ XES measured at the LCLS X-ray free-electron laser.

Transient X-ray emission spectroscopy

$K\alpha$ and $K\beta$ XES monitors the 2p-to-1s and 3p-to-1s fluorescence, respectively, following 1s core-hole X-ray ionization. The strong influence of the exchange interaction between the $n\text{p}$ and the 3d valence electrons make the XES spectra sensitive to the total spin multiplicity on the iron center.^{56,61–64} The spin-sensitivity of the two techniques is illustrated by the reference spectra shown in Fig. 4A and B. The reference spectra are measured for a series of Fe-centered compounds with ground state electronic structure similar to those of the potential excited states of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$. By selecting reference compounds with ligand bond covalency similar to $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$, good agreement between the reference spectra and the measured data is expected. By subtracting the reference signal of the ground state compound, from the references for the potential excited states, the reference difference spectra, illustrated in Fig. 4C and D can be constructed.

The $K\beta$ spectra exhibit the clearest spectral shape variation between different electronic states, such that direct comparison between transient $K\beta$ data and reference difference spectra allows for robust identification of the excited electronic states. Meanwhile, the $K\alpha$ emission is roughly an order of magnitude more intense than the $K\beta$ emission, which makes the $K\alpha$ data well suited to derive the time scales of the underlying relaxation dynamics. Based on these considerations, we have measured $K\alpha$ single-energy kinetics to determine the characteristic time scales of the experiment (time-zero, instrument-response function (IRF), and lifetimes), and measured full $K\beta$ transient spectra for robust excited state identification. Since the $K\alpha$ and $K\beta$ data were recorded simultaneously, the time zero, temporal resolution, and quantum yields of the two datasets are identical. Therefore, the characteristic time constants of the experimental parameters, as well as the excited state lifetimes, determined from the high signal-to-noise $K\alpha$ kinetic data can be used in a global analysis framework for the $K\beta$ spectral data.

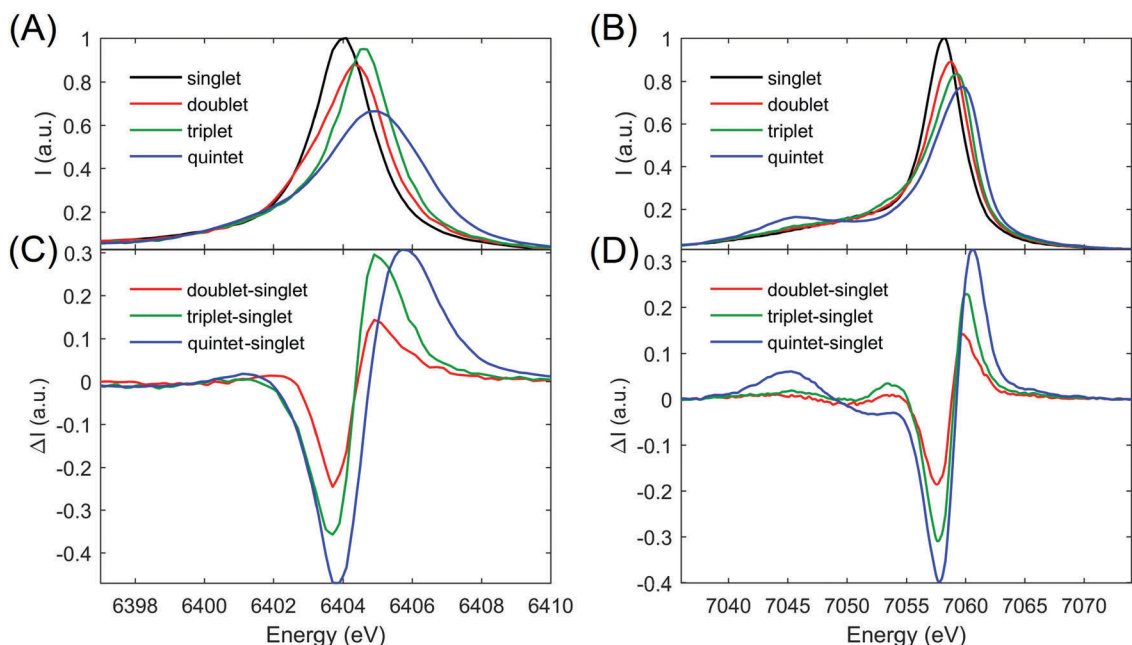


Fig. 4 $K\alpha_1$ (A) and $K\beta$ (B) reference spectra measured for a series of Fe model complexes of similar spin state configuration as the potential excited states in this study. $K\alpha$ references are $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ (singlet), $[\text{Fe}(\text{CN})_6]^{3-}$ (doublet), $\text{Fe}(\text{phthalocyanine})$ (triplet), and $[\text{Fe}(1,10\text{-phenanthroline})_2(\text{NCS})_2]$ (quintet). $K\beta$ references are taken from Zhang *et al.*¹ Reference difference spectra (C and D) constructed from the $K\alpha$ and $K\beta$ references generated by subtracting the ground state reference from each of the excited state references.

The $K\alpha$ kinetics shown in Fig. 5 were recorded at 6404 eV. As reported elsewhere,⁶² and seen from the reference difference spectra (Fig. 4C), the $K\alpha$ difference intensity at 6404 eV allow us to monitor the total spin increase on the Fe center. We observe that the growth of the transient signal is delayed with respect to a broadened Heaviside step function. This delay arises from an

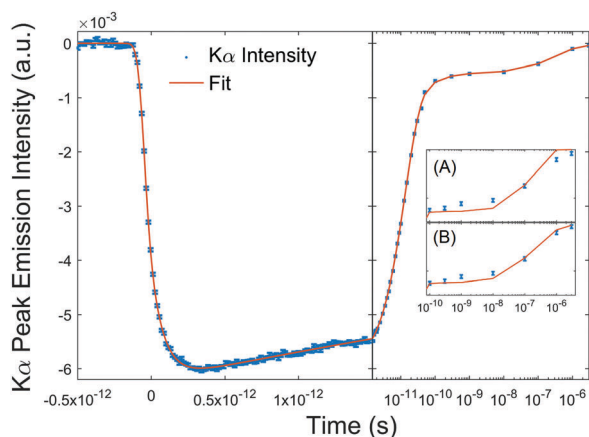


Fig. 5 Kinetics of the $K\alpha_1$ peak emission intensity measured at 6404 eV following MLCT excitation of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in water. Blue data points are binned in 10 fs steps and supplied with error bars indicating the standard error. The red curve shows the fit of the data as described in the main text. Insets (A) and (B) show fits to the late data points using first and second order reaction kinetics, respectively. The second order reaction kinetics (B), capture the evolution of the data better, illustrating that the kinetics are due to charge-recombination of photoionized $[\text{Fe}(\text{bpy})(\text{CN})_4]^-$ with solvated electrons, as described in detail in the ESI.†

excited state relaxation that results in a secondary increase in the Fe spin-moment, delayed with respect to excitation of the MLCT state. Overall, the growth is best described by the sum of a Heaviside step function and exponential grow-in broadened by a Gaussian representing the instrument-response function. The step function accounts for the instantaneously populated $^1\text{MLCT}$ and photoionization products.^{65,66} These species each have a single unpaired electron on the Fe center, making them formal doublets with respect to the Fe center, and therefore they cannot be distinguished with the XES data. The exponential growth accounts for the transition from the MLCT to the secondary excited state of higher spin-moment. The fact that the Fe spin-moment increases for the secondary excited state confirms the assumption, based on the optical TA data, that the decay of the MLCT state leads to the population of an MC state with higher spin. The decay of the transient $K\alpha$ signal is dominated by the single-exponential lifetime of the secondary MC excited state, but also includes a weaker long time-scale component for the recovery of the photoionization products.

In fitting the data to the kinetic model described above, a time-zero and IRF (rms) of -54 ± 4 fs and 34 ± 2 fs are recovered, with the uncertainty reflecting the standard deviation. The exponential component of the growth accounts for a $40 \pm 4\%$ increase of the total signal and has a time constant of 87 ± 5 fs, which we assign to the MLCT-to-MC transition. The decay is dominated by a component with a 13.1 ± 0.4 ps lifetime, readily assigned to the decay of the MC state and recovery of the ground state accounting for 84% of the decay. The remaining decay of the transient signal occurs on the hundreds-of-nanoseconds to microsecond time scale. The MLCT-to-MC

transition observed by the XES is faster than the lifetime associated with SAS1 in the TA analysis. This discrepancy likely reflects the challenge of differentiating intramolecular vibrational relaxation and population decay in TA, combined with the challenges of properly subtracting out cross-phase modulation artefacts.

The long time scale dynamics of the ground state recovery matches the expected second-order kinetics (as described in the ESI[†]) for the recovery from the photoionized $[\text{Fe}(\text{bpy})(\text{CN})_4]^-$.⁶⁰ Comparative fits to the long time delay data using first- and second order rate expressions for the decay are illustrated in Fig. 5 as insets (A) and (B), respectively. A detailed discussion of the mechanism and time scale of the recovery can be found in the ESI.[†] The increased ratio of photoionization product (16%) compared to the TA data is due to the increased laser power used for the XES experiments.

In summary, the $K\alpha$ decay kinetics can be described by the same model used to interpret the TA data. Since the TA experiment shows that the dynamics of the single-photon excited state pathway is independent on the amount of two-photon excitation, the XES data can be used to identify the spin state of the 13 ps MC intermediate associated with SAS2 in the TA analysis. For this excited state identification, we rely on the $K\beta$ data recorded simultaneously with the $K\alpha$ data.

Transient $K\beta$ differential spectra measured at multiple time delays are presented in Fig. 6A. Qualitative comparison of the data recorded at 0.5–5 ps (delays dominated by the MC intermediate) with the reference spectra in Fig. 4D reveals that the $K\beta'$ shoulder at 7045 eV is very weak. This allows the high spin ^5MC state to be eliminated as the dominant product of the MLCT decay and makes a ^3MC the most likely candidate for the MC excited state. To explicitly identify the nature of the MC state and derive excitation and speciation fractions, we have conducted a global analysis of the transient $K\beta$ data using a three-state model that includes the MLCT state, the MC state,

and the $[\text{Fe}(\text{bpy})(\text{CN})_4]^-$ photoproduct. Using the excited state cascade and time scales from the $K\alpha$ analysis, the species associated spectra (SAS) for the excited states were extracted.

The extracted SAS are shown in Fig. 6B. SAS1 describes the signal that is lost with the 85 fs time constant after the excitation. SAS2 describes the signal that grows in as SAS1 decays, and has a lifetime of 13.1 ps. SAS1 was definitively identified as an MLCT state from the optical TA, and could therefore be used to establish a scaling factor between the SAS and reference difference spectra by minimizing the chi-squared between SAS1 and the MLCT (doublet–singlet) reference difference signal. Upon establishing this scaling factor, a quantitative comparison of SAS2 (representing the MC excited state) can be made against the triplet and quintet reference difference spectra (all plotted in Fig. 6B). It is immediately apparent that the signal shape, and in particular signal amplitude, of SAS2 matches the reference spectrum obtained for the ^3MC state. Taken together with the lack of additional fast time constants in the excited state cascade, the agreement between SAS2 and the ^3MC reference difference spectrum definitively assigns the ultrafast interconversion observed in the optical TA and $K\alpha$ data to an MLCT-to- ^3MC interconversion. Furthermore, the spectral features of SAS2 in the optical TA analysis can now be assigned as characteristic of the ^3MC state. The ultrafast XES measurements provide a robust assignment of the ^3MC state, which enables the interpretation of the ESA seen in the optical pump–probe measurement. The most significant feature of the spectrum is the excited state absorption feature at 570 nm, to the red side of the ground state bleach. Similar, but somewhat weaker, ESA features have been observed for ^3MC states assigned of Ru-centered polypyridyl complexes.⁶⁷

Solvent-dependent electronic structure

The solvent Lewis acidity strongly influences the excited state dynamics of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$. Increasing solvent Lewis acidity

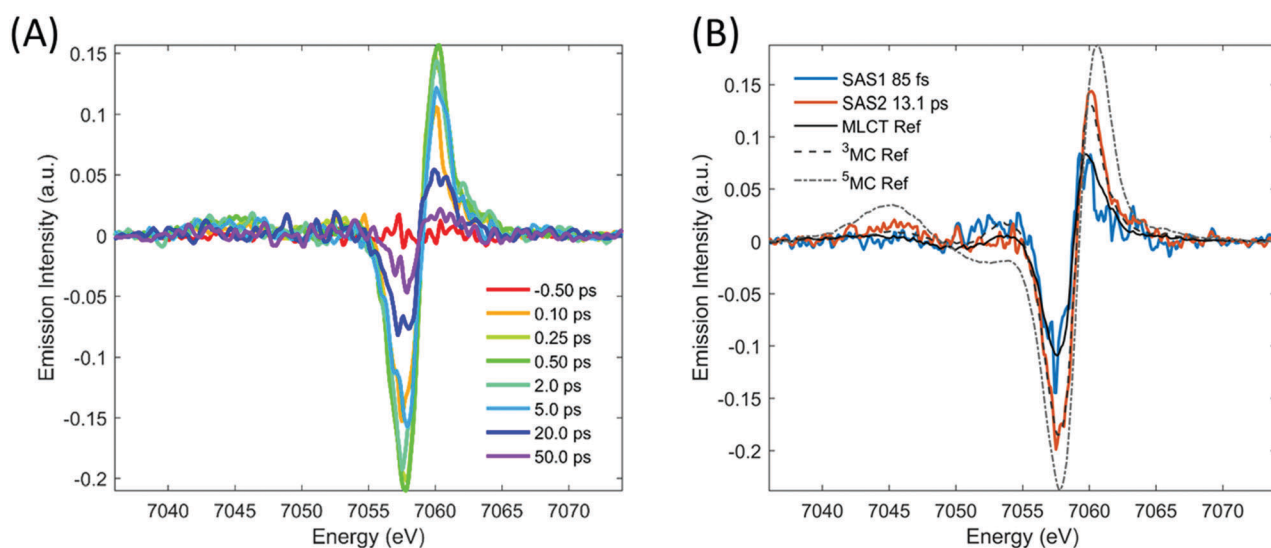


Fig. 6 Transient $K\beta$ spectra and analysis. (A) Transient $K\beta$ emission signal of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in water. (B) Species associated spectra (SAS, red and blue) from a global analysis, invoking the model used for fitting the TA and $K\alpha$ kinetics and the reference difference spectra of Fig. 4D scaled by mapping the MLCT reference to SAS1.

leads to rapid excited state interconversion of the otherwise relatively long lived MLCT state. Furthermore, the ^3MC intermediate, which is expected to govern the MLCT decay, is stabilized to the point where it can be considered a well-isolated intermediate in the non-radiative relaxation. These changes in excited state dynamics reflect the solvent-dependent changes in the electronic structure of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$. A qualitative molecular orbital (MO) diagram of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in weak Lewis acid solvent is shown in Fig. 7A. Both t_{2g} and e_g levels of the cyano-pyridyl complex increase upon substitution of bpy with CN^- ligands, as demonstrated by the reduction in the metal oxidation potential, the reduction in the MLCT excitation energy, and the elimination of the quintet state from the excited state relaxation pathway.^{1,2,39,44} These energetic changes in the redox potentials are readily explained by the σ and π^* energy levels of the CN^- ligand being of higher than the σ and π^* levels of bpy, as illustrated schematically in Fig. 7A. This MO picture is consistent with the projected potential energy surfaces (PES), calculated for $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in DMSO, which are illustrated in a schematic form in Fig. 7B, based on the calculations in Zhang *et al.*¹

Increasing solvent Lewis acidity leads to a blue-shift of the MLCT absorption bands, as shown in Fig. 1. A similar solvation response has also been observed for Ru-centered cyano-pyridyls, where the solvent-dependent blue-shift has been shown to increase linearly with both the number of CN^- ligands and the Lewis acidity of the solvent.³⁹ The blue-shift is interpreted as a result of the high Lewis acidity solvent shifting electron density of

the CN^- orbitals away from the metal center, decreasing the $\pi(\text{CN}^-) \rightarrow d(\text{Fe})$ π -bonding and increasing the $d(\text{Fe}) \rightarrow \pi^*(\text{CN}^-)$ π -backbonding. Both of these changes stabilize the Fe t_{2g} levels relative to the (poly)pyridyl centered π^* levels (which are relatively unaffected by solvation).

The influence of lowering the t_{2g} levels on the PES is illustrated in Fig. 7C. The GS having 6 t_{2g} electrons is stabilized more than the MLCT and the ^3MC states each having 5 t_{2g} electrons, which are in turn stabilized more than the ^5MC state having only 4 t_{2g} electrons. Since the GS is stabilized the most, all excited states undergo an apparent destabilization since the PES in Fig. 7C are normalized relative to the ground state energy. Since the ^5MC state has two less t_{2g} electrons than the GS, and the ^3MC and MLCT states have one less t_{2g} electron than the GS, the relative destabilization of the ^5MC state is twice that of the ^3MC and MLCT states.

While the changes in the t_{2g} energy levels can be addressed by steady state measurements, the changes in e_g levels are most clearly addressed with transient measurements since the e_g levels only influence the metal centered electronic excited states. To explain the significant shortening of the MLCT lifetime of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in water relative to DMSO, the ^3MC state has to intersect the MLCT surface closer to its energy minimum. Therefore, the ^3MC state must be stabilized relative to the MLCT state, which indicates that the e_g levels also decrease in energy with increasing solvent Lewis acidity. This stabilization of the σ -bonded e_g orbitals is rationalized by the

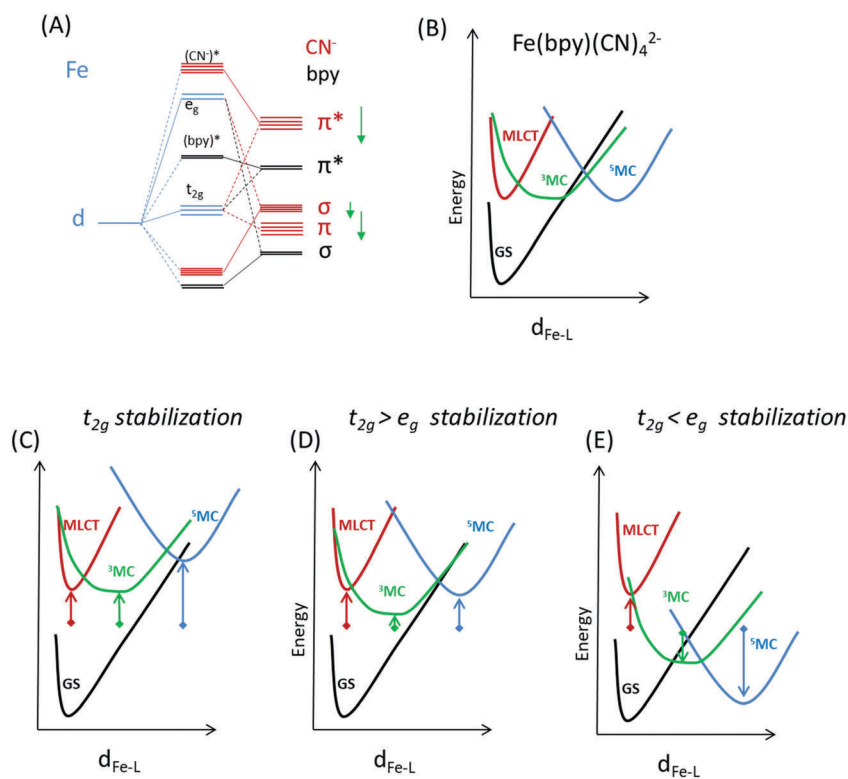


Fig. 7 (A) Schematic molecular orbital diagram for $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$. Green arrows indicate the shift in energy levels upon solvation in high Lewis acidity solvents. (B) Schematic of the potential energy surfaces of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in low Lewis acidity solvents.¹ (C–E) Effect of variable t_{2g} and e_g orbital stabilization on the MLCT and MC excited state PES.

same argument for the π -bonded t_{2g} orbital stabilization discussed above. Stabilization of the e_g orbitals will shift down the ^3MC and ^5MC states relative to the illustration in Fig. 7C, which only accounts for the t_{2g} orbital stabilization. Two cases accounting for both t_{2g} and e_g stabilization are illustrated in Fig. 7D and E.

Stabilization of the e_g orbitals is consistent with the solvent dependence of the MLCT lifetimes in Ru-centered cyanopyridyls.⁴⁴ For these Ru complexes, the relationship between MLCT lifetime and solvent acceptor number (a measure of the solvent Lewis acidity) could be described by assuming that the ^3MC energy was unchanged relative to the GS, requiring identical stabilization of the t_{2g} and the e_g levels. Interestingly, our experiments show that the ^3MC lifetime, which is too short to be clearly observed in low Lewis acidity solvents,¹ is significantly extended in high Lewis acidity solvents, meaning that a reaction barrier for the inter-system crossing and internal conversion back to the GS is required. This reaction barrier can only arise by stabilizing the e_g levels by a smaller amount than the t_{2g} levels, such that the MC states are destabilized relative to the GS (Fig. 7D). If the e_g levels are destabilized equal to or more than the t_{2g} levels, then the MC states are stabilized with respect to the GS, as illustrated in Fig. 7E. In this case, it becomes impossible to arrange the potential energy surfaces such that there is a reaction barrier to the ground state, while maintaining a significant reaction barrier to the ^5MC state consistent with the stabilization of the ^3MC and ^5MC states being proportionally with their number of e_g electrons. Since we observe no signature of the ^5MC state in the excited state cascade, we conclude the ordering of the potential energy surfaces of $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$ in high Lewis acidity solvents match the illustration in Fig. 7D. In summary, the observation that the MLCT lifetime decreases with increasing solvent Lewis acidity shows that the e_g levels are stabilized in strong Lewis acid solvents, but not as strongly as the t_{2g} levels.

To support our interpretation of the solvent influence on the excited state energy levels, we have performed preliminary DFT calculations with a polarizable continuum description of both water and DMSO, and by including explicit water molecules at the N lone-pair site of the CN^- groups. The effects of implicit and explicit solvent interactions on the excited state energy levels are summarized in Table S1 in the ESI.[†] It is observed that there is virtually no difference between the energy levels calculated for water and DMSO when representing the solvents with continuum models. However, upon inclusion of the explicit waters at the N lone pair sites of the CN^- groups, the DFT calculations shows a destabilization of $^3\text{MLCT}$, ^5MC , and ^3MC states with respect to the ground state, in parallel to our experimental observations. In accordance with our interpretation of the experimental results, the calculations also show that the $^3\text{MLCT}$ state is destabilized the most, and the ^3MC state is destabilized the least. The minimum energy structures of our calculations (presented in Fig. S5, ESI[†]) shows that the water orients to form hydrogen bonds between to the cyanide N lone-pairs, supporting our interpretation of the experimental results that the reordering of the excited state level arises from explicit electronic interactions between solute and solvent.

Closing remarks

Combined ultrafast optical and XES measurements provide a clear picture for the MLCT excited state relaxation dynamics in $[\text{Fe}(\text{bpy})(\text{CN})_4]^{2-}$. In high Lewis acidity solvents, coordination of the CN^- ligands to the solvent significantly changes the lifetimes of the MLCT and the ^3MC excited states relative to their lifetimes in low Lewis acidity solvents. The influence of the solvent on the excited state relaxation dynamics arise from inequivalent solvent stabilization of the CN^- and bpy energy levels, shifting the CN^- -influenced t_{2g} and e_g orbitals down relative to the bpy π^* levels, which in turn modifies the relative energies of the MLCT, ^3MC , and ^5MC excited state potential energy surfaces. Our study presents a clear identification and characterization of a metastable Fe-centered ^3MC state in a hexacoordinated Fe centered complex.

Extracting this level of detail about the excited state relaxation mechanism has been made possible by the combined application of time resolved UV-visible TA and XES. Combining these techniques has allowed us to address not only the MLCT state dynamics, but also to identify the role of MC excited states in the electronic excited state relaxation, obtaining a more detailed picture of the excited state potential energy surfaces involved in the relaxation process. The increased robustness of the relaxation mechanisms extracted from the combined analysis of XES and TA measurements make them more amenable targets for computational quantum dynamics studies. Moreover, the relative simplicity of the Fe cyano-pyridyl compounds, and the fact that their potential energy landscape can be extensively modified by the choice of solvent, enabled a systematic exploration of the relationship between the potential energy landscape and the excited state lifetime and relaxation pathways.

These experiments also act as a stepping stone towards the investigation of Fe-centered photoactive molecular systems. With regards to future XES measurements, the good agreement with our reference spectra and the difference signal extracted from global analysis, provide further support for using the XES technique to identify the excited state species in 3d transition metal centered systems. With regards to future TA measurements, the clearly identified and isolated ^3MC state presented here exhibit significant excited state absorption features throughout the visible spectrum. This should significantly help the assignment of excited state absorptions in future TA measurements on chemically related Fe coordination complexes.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 W. Zhang, *et al.*, Manipulating charge transfer excited state relaxation and spin crossover in iron coordination complexes with ligand substitution, *Chem. Sci.*, 2016, **8**, 515–523.
- 2 Y. Liu, *et al.*, A Heteroleptic Ferrous Complex with Mesoionic Bis(1,2,3-triazol-5-ylidene) Ligands: Taming the MLCT Excited State of Iron(II), *Chem. – Eur. J.*, 2015, **21**, 3628–3639.
- 3 J. He, B. Wang, S. Chang and T. Chen, Ruthenium-Based Photosensitizers for Dye-Sensitized Solar Cells, in *Organometallics and Related Molecules for Energy Conversion*, ed. W.-Y. Wong, Springer, Berlin Heidelberg, 2015, pp. 91–114, DOI: 10.1007/978-3-662-46054-2_4.
- 4 E. Baranoff, First-Row Transition Metal Complexes for the Conversion of Light into Electricity and Electricity into Light, in *Organometallics and Related Molecules for Energy Conversion*, ed. W.-Y. Wong, Springer, Berlin Heidelberg, 2015, pp. 61–90, DOI: 10.1007/978-3-662-46054-2_3.
- 5 J. V. Caspar and T. J. Meyer, Photochemistry of tris(2,2'-bipyridine)ruthenium(2+) ion ($\text{Ru}(\text{bpy})_3^{2+}$). Solvent effects, *J. Am. Chem. Soc.*, 1983, **105**, 5583–5590.
- 6 A. Juris, *et al.*, Ru(II) polypyridine complexes: photophysics, photochemistry, electrochemistry, and chemiluminescence, *Coord. Chem. Rev.*, 1988, **84**, 85–277.
- 7 N. H. Damrauer, *et al.*, Femtosecond Dynamics of Excited-State Evolution in $[\text{Ru}(\text{bpy})_3]^{2+}$, *Science*, 1997, **275**, 54–57.
- 8 M. Abrahamsson, *et al.*, A 3.0 μs Room Temperature Excited State Lifetime of a Bistridentate RuII–Polypyridine Complex for Rod-like Molecular Arrays, *J. Am. Chem. Soc.*, 2006, **128**, 12616–12617.
- 9 D. G. Brown, N. Sanguantrakun, B. Schulze, U. S. Schubert and C. P. Berlinguette, Bis(tridentate) Ruthenium–Terpyridine Complexes Featuring Microsecond Excited-State Lifetimes, *J. Am. Chem. Soc.*, 2012, **134**, 12354–12357.
- 10 S. Nozawa, *et al.*, Direct Probing of Spin State Dynamics Coupled with Electronic and Structural Modifications by Picosecond Time-Resolved XAFS, *J. Am. Chem. Soc.*, 2010, **132**, 61–63.
- 11 J. E. Monat and J. K. McCusker, Femtosecond Excited-State Dynamics of an Iron(II) Polypyridyl Solar Cell Sensitizer Model, *J. Am. Chem. Soc.*, 2000, **122**, 4092–4097.
- 12 H. T. Lemke, *et al.*, Femtosecond X-ray Absorption Spectroscopy at a Hard X-ray Free Electron Laser: Application to Spin Crossover Dynamics, *J. Phys. Chem. A*, 2013, **117**, 735–740.
- 13 N. Huse, *et al.*, Femtosecond Soft X-ray Spectroscopy of Solvated Transition-Metal Complexes: Deciphering the Interplay of Electronic and Structural Dynamics, *J. Phys. Chem. Lett.*, 2011, **2**, 880–884.
- 14 A. Cannizzo, *et al.*, Light-induced spin crossover in Fe(II)-based complexes: The full photocycle unraveled by ultrafast optical and X-ray spectroscopies, *Coord. Chem. Rev.*, 2010, **254**, 2677–2686.
- 15 K. Haldrup, *et al.*, Observing Solvation Dynamics with Simultaneous Femtosecond X-ray Emission Spectroscopy and X-ray Scattering, *J. Phys. Chem. B*, 2016, **120**, 1158–1168.
- 16 N. Huse, *et al.*, Photo-Induced Spin-State Conversion in Solvated Transition Metal Complexes Probed via Time-Resolved Soft X-ray Spectroscopy, *J. Am. Chem. Soc.*, 2010, **132**, 6809–6816.
- 17 M. Khalil, *et al.*, Picosecond X-ray Absorption Spectroscopy of a Photoinduced Iron(II) Spin Crossover Reaction in Solution, *J. Phys. Chem. A*, 2006, **110**, 38–44.
- 18 G. Auböck and M. Chergui, Sub-50 fs photoinduced spin crossover in $[\text{Fe}(\text{bpy})_3]^{2+}$, *Nat. Chem.*, 2015, **7**, 629–633.
- 19 J. K. McCusker, *et al.*, Subpicosecond 1MLCT.fwdarw. 5T2 intersystem crossing of low-spin polypyridyl ferrous complexes, *J. Am. Chem. Soc.*, 1993, **115**, 298–307.
- 20 A. Marino, *et al.*, The Role of Ligand-Field States in the Ultrafast Photophysical Cycle of the Prototypical Iron(II) Spin-Crossover Compound $[\text{Fe}(\text{ptz})_6](\text{BF}_4)_2$, *Angew. Chem., Int. Ed.*, 2014, **53**, 3863–3867.
- 21 W. Zhang, *et al.*, Tracking excited-state charge and spin dynamics in iron coordination complexes, *Nature*, 2014, **509**, 345–348.
- 22 C. Sousa, *et al.*, Ultrafast Deactivation Mechanism of the Excited Singlet in the Light-Induced Spin Crossover of $[\text{Fe}(2,2'\text{-bipyridine})_3]^{2+}$, *Chem. – Eur. J.*, 2013, **19**, 17541–17551.
- 23 W. Gawelda, *et al.*, Ultrafast Nonadiabatic Dynamics of $[\text{FeII}(\text{bpy})_3]^{2+}$ in Solution, *J. Am. Chem. Soc.*, 2007, **129**, 8199–8206.
- 24 C. Consani, *et al.*, Vibrational Coherences and Relaxation in the High-Spin State of Aqueous $[\text{FeII}(\text{bpy})_3]^{2+}$, *Angew. Chem., Int. Ed.*, 2009, **48**, 7184–7187.

- 25 Y. Liu, *et al.*, Towards longer-lived metal-to-ligand charge transfer states of iron(II) complexes: an N-heterocyclic carbene approach, *Chem. Commun.*, 2013, **49**, 6412–6414.
- 26 L. A. Fredin, *et al.*, Exceptional Excited-State Lifetime of an Iron(II)–N-Heterocyclic Carbene Complex Explained, *J. Phys. Chem. Lett.*, 2014, **5**, 2066–2071.
- 27 T. C. B. Harlang, *et al.*, Iron sensitizer converts light to electrons with 92% yield, *Nat. Chem.*, 2015, **7**, 883–889.
- 28 T. Duchanois, *et al.*, An Iron-Based Photosensitizer with Extended Excited-State Lifetime: Photophysical and Photo-voltaic Properties, *Eur. J. Inorg. Chem.*, 2015, 2469–2477.
- 29 M. Pápai, G. Vankó, T. Rozgonyi and T. J. Penfold, High-Efficiency Iron Photosensitizer Explained with Quantum Wavepacket Dynamics, *J. Phys. Chem. Lett.*, 2016, **7**, 2009–2014.
- 30 Y. Liu, P. Persson, V. Sundström and K. Wärnmark, Fe N-Heterocyclic Carbene Complexes as Promising Photosensitizers, *Acc. Chem. Res.*, 2016, **49**, 1477–1485.
- 31 L. Liu, *et al.*, A new record excited state 3MLCT lifetime for metal organic iron(II) complexes, *Phys. Chem. Chem. Phys.*, 2016, **18**, 12550–12556.
- 32 M. Reinhard, *et al.*, Photoaquation Mechanism of Hexacyanoferrate(II) Ions: Ultrafast 2D UV and Transient Visible and IR Spectroscopies, *J. Am. Chem. Soc.*, 2017, **139**, 7335–7347.
- 33 C. D. Graaf and C. Sousa, On the role of the metal-to-ligand charge transfer states in the light-induced spin crossover in FeII (bpy)₃, *Int. J. Quantum Chem.*, 2011, **111**, 3385–3393.
- 34 A. Moguilevski, *et al.*, Ultrafast Spin Crossover in [FeII(bpy)₃]²⁺: Revealing Two Competing Mechanisms by Extreme Ultraviolet Photoemission Spectroscopy, *ChemPhysChem*, 2017, **18**, 465–469.
- 35 I. M. Dixon, F. Alary, M. Boggio-Pasqua and J.-L. Heully, Reversing the relative 3MLCT–3MC order in Fe(II) complexes using cyclometallating ligands: a computational study aiming at luminescent Fe(II) complexes, *Dalton Trans.*, 2015, **44**, 13498–13503.
- 36 L. L. Jamula, A. M. Brown, D. Guo and J. K. McCusker, Synthesis and Characterization of a High-Symmetry Ferrous Polypyridyl Complex: Approaching the 5T₂/3T₁ Crossing Point for FeII, *Inorg. Chem.*, 2014, **53**, 15–17.
- 37 I. M. Dixon, F. Alary, M. Boggio-Pasqua and J.-L. Heully, The (N4C2)₂– Donor Set as Promising Motif for Bis(tridentate) Iron(II) Photoactive Compounds, *Inorg. Chem.*, 2013, **52**, 13369–13374.
- 38 M. Pápai, T. J. Penfold and K. B. Møller, Effect of *tert*-Butyl Functionalization on the Photoexcited Decay of a Fe(II)–N-Heterocyclic Carbene Complex, *J. Phys. Chem. C*, 2016, **120**, 17234–17241.
- 39 C. J. Timpson, C. A. Bignozzi, B. P. Sullivan, E. M. Kober and T. J. Meyer, Influence of Solvent on the Spectroscopic Properties of Cyano Complexes of Ruthenium(II), *J. Phys. Chem.*, 1996, **100**, 2915–2925.
- 40 I. Georgieva, A. J. A. Aquino, N. Trendafilova, P. S. Santos and H. Lischka, Solvatochromic and Ionochromic Effects of Iron(II)bis(1,10-phenanthroline)dicyano: a Theoretical Study, *Inorg. Chem.*, 2010, **49**, 1634–1646.
- 41 W. Linert and V. Gutmann, Structural and electronic responses of coordination compounds to changes in the molecule and molecular environment, *Coord. Chem. Rev.*, 1992, **117**, 159–183.
- 42 A. Al-alousy and J. Burgess, Bis(1,10-phenanthroline)dicyanoiron(II): an almost universal inorganic solvent polarity indicator, *Inorg. Chim. Acta*, 1990, **169**, 167–170.
- 43 L. Fodor, G. Lendvay and A. Horváth, Solvent Dependence of Absorption and Emission Spectra of Ru(bpy)₂(CN)₂: Experiment and Explanation Based on Electronic Structure Theory, *J. Phys. Chem. A*, 2007, **111**, 12891–12900.
- 44 M. T. Indelli, C. A. Bignozzi, F. Scandola and J.-P. Collin, Design of Long-Lived Ru(II) Terpyridine MLCT States. Tricyano Terpyridine Complexes, *Inorg. Chem.*, 1998, **37**, 6084–6089.
- 45 J. R. Winkler, C. Creutz and N. Sutin, Solvent tuning of the excited-state properties of (2,2′-bipyridine)tetracyanoferrate(II): direct observation of a metal-to-ligand charge-transfer excited state of iron(II), *J. Am. Chem. Soc.*, 1987, **109**, 3470–3471.
- 46 T. J. Penfold, *et al.*, X-ray Spectroscopic Study of Solvent Effects on the Ferrous and Ferric Hexacyanide Anions, *J. Phys. Chem. A*, 2014, **118**, 9411–9418.
- 47 Note that while the [Fe(bpy)(CN)₄]^{2–} is formally of C_{2v} symmetry we refer to the d-π and d-σ* orbitals with the octahedral eg and t_{2g} notation to facilitate direct comparison between the different systems that are all near octahedral symmetry.
- 48 K. Haldrup, *et al.*, Guest–Host Interactions Investigated by Time-Resolved X-ray Spectroscopies and Scattering at MHz Rates: Solvation Dynamics and Photoinduced Spin Transition in Aqueous Fe(bipy)₃²⁺, *J. Phys. Chem. A*, 2012, **116**, 9878–9887.
- 49 S. E. Canton, *et al.*, Visualizing the non-equilibrium dynamics of photoinduced intramolecular electron transfer with femto-second X-ray pulses, *Nat. Commun.*, 2015, **6**, 6359.
- 50 G. Vankó, *et al.*, Detailed Characterization of a Nanosecond-Lived Excited State: X-ray and Theoretical Investigation of the Quintet State in Photoexcited [Fe(terpy)₂]²⁺, *J. Phys. Chem. C*, 2015, **119**, 5888–5902.
- 51 Z. Fox, *et al.*, Monitoring Excited State Charge Transfer of Transition Metal Mixed-Valence Complexes with Femto-second X-ray Absorption and Emission Spectroscopy, International Conference on Ultrafast Phenomena (2016), UTh4A.2 (Optical Society of America, 2016), DOI: 10.1364/UP.2016.UTh4A.2.
- 52 R. Alonso-Mori, *et al.*, Photon-in photon-out hard X-ray spectroscopy at the Linac Coherent Light Source, *J. Synchrotron Radiat.*, 2015, **22**, 612–620.
- 53 M. Chollet, *et al.*, The X-ray Pump–Probe instrument at the Linac Coherent Light Source, *J. Synchrotron Radiat.*, 2015, **22**, 503–507.
- 54 H. T. Philipp, M. W. Tate and S. M. Gruner, Low-flux measurements with Cornell’s LCLS integrating pixel array detector, *J. Instrum.*, 2011, **6**, C11006.
- 55 R. Alonso-Mori, *et al.*, A multi-crystal wavelength dispersive x-ray spectrometer, *Rev. Sci. Instrum.*, 2012, **83**, 073114.

- 56 P. Glatzel and U. Bergmann, High resolution 1s core hole X-ray spectroscopy in 3d transition metal complexes – electronic and structural information, *Coord. Chem. Rev.*, 2005, **249**, 65–95.
- 57 H. E. Toma and M. S. Takasugi, Spectroscopic studies of preferential and asymmetric solvation in substituted cyanoiron(II) complexes, *J. Solution Chem.*, 1983, **12**, 547–561.
- 58 A. M. Brown, C. E. McCusker and J. K. McCusker, Spectro-electrochemical identification of charge-transfer excited states in transition metal-based polypyridyl complexes, *Dalton Trans.*, 2014, **43**, 17635–17646.
- 59 K. S. Kjær, *et al.*, Ligand manipulation of charge transfer excited state relaxation and spin crossover in [Fe(2,2'-bipyridine)2(CN)2], *Struct. Dyn.*, 2017, **4**, 044030.
- 60 A. Horváth, J. Szöke and L. Wojnárovits, Mechanism of photoinduced redox reactions in aqueous solutions of [Fe(bpy)(CN)4]2[−], *Inorg. Chim. Acta*, 1991, **179**, 97–104.
- 61 F. de Groot, High-Resolution X-ray Emission and X-ray Absorption Spectroscopy, *Chem. Rev.*, 2001, **101**, 1779–1808.
- 62 G. Vankó, *et al.*, Probing the 3d Spin Momentum with X-ray Emission Spectroscopy: The Case of Molecular-Spin Transitions, *J. Phys. Chem. B*, 2006, **110**, 11647–11653.
- 63 N. Lee, T. Petrenko, U. Bergmann, F. Neese and S. DeBeer, Probing Valence Orbital Composition with Iron Kβ X-ray Emission Spectroscopy, *J. Am. Chem. Soc.*, 2010, **132**, 9715–9727.
- 64 G. Vankó, *et al.*, Spin-state studies with XES and RIXS: From static to ultrafast, *J. Electron Spectrosc. Relat. Phenom.*, 2013, **188**, 166–171.
- 65 N. A. Anderson, K. Hang, J. B. Asbury and T. Lian, Ultrafast mid-IR detection of the direct precursor to the presolvated electron following electron ejection from ferrocyanide, *Chem. Phys. Lett.*, 2000, **329**, 386–392.
- 66 M. Reinhard, *et al.*, Photooxidation and photoaquation of iron hexacyanide in aqueous solution: A picosecond X-ray absorption study, *Struct. Dyn.*, 2014, **1**, 024901.
- 67 Q. Sun, *et al.*, On the role of ligand-field states for the photophysical properties of ruthenium(II) polypyridyl complexes, *Coord. Chem. Rev.*, 2015, **282–283**, 87–99.